

# ***HUMIC ACID***

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**SYNTHETIC SOIL-EXTRACT MATERIALS AND  
MEDICAMENTS FOR INFLUENZA VIRUSES  
BASED THEREON**

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(54) **SYNTHETIC SOIL-EXTRACT MATERIALS AND MEDICAMENTS FOR INFLUENZA VIRUSES BASED THEREON**

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(73) Assignee: **Laub Biochemicals Corp.**, Newport Beach, CA (US)

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(\* ) Notice: Subject to any disclaimer, the term of this patent is extended or adjusted under 35 U.S.C. 154(b) by 0 days.

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This patent is subject to a terminal disclaimer.

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**Related U.S. Application Data**

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(60) Provisional application No. 60/297,669, filed on Jun. 12, 2001.

(51) **Int. Cl.**<sup>7</sup> ..... **A61K 31/74**; A61K 31/765; A61K 31/35

(52) **U.S. Cl.** ..... **424/78.02**; 424/48.37; 514/456

(58) **Field of Search** ..... 434/78.02, 78.37; 514/456

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*Primary Examiner*—Thurman K. Page

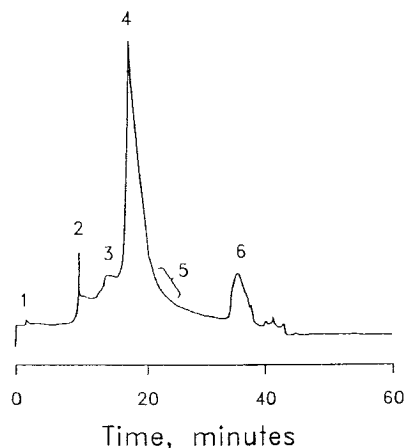
*Assistant Examiner*—Blessing Fubara

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(57) **ABSTRACT**

Phenolic polymers are prepared by oxidizing and polymerizing starting organic compounds comprising at least one hydroxyl group and at least one carbonyl group or at least two hydroxyl groups on an aromatic structure. One or more inorganic compounds or salts is added and the solution is allowed to stand at about 20° C. to 80° C. for a period of about at least 2 hours. Salt molecules as well as starting compounds and other low molecular-weight materials below about 500 to about 10,000 daltons are removed from the product solution. Purified phenolic polymers are prepared in concentrated aqueous solution or in dried powder form in a final step, if necessary. The resultant phenolic polymers exhibit physicochemical properties strongly resembling those of typical commercially-available natural-product soil extracts. The materials are active influenza anti-viral agents, and are effective in anti-viral compositions for treating or preventing influenza viral diseases.

**38 Claims, 2 Drawing Sheets**

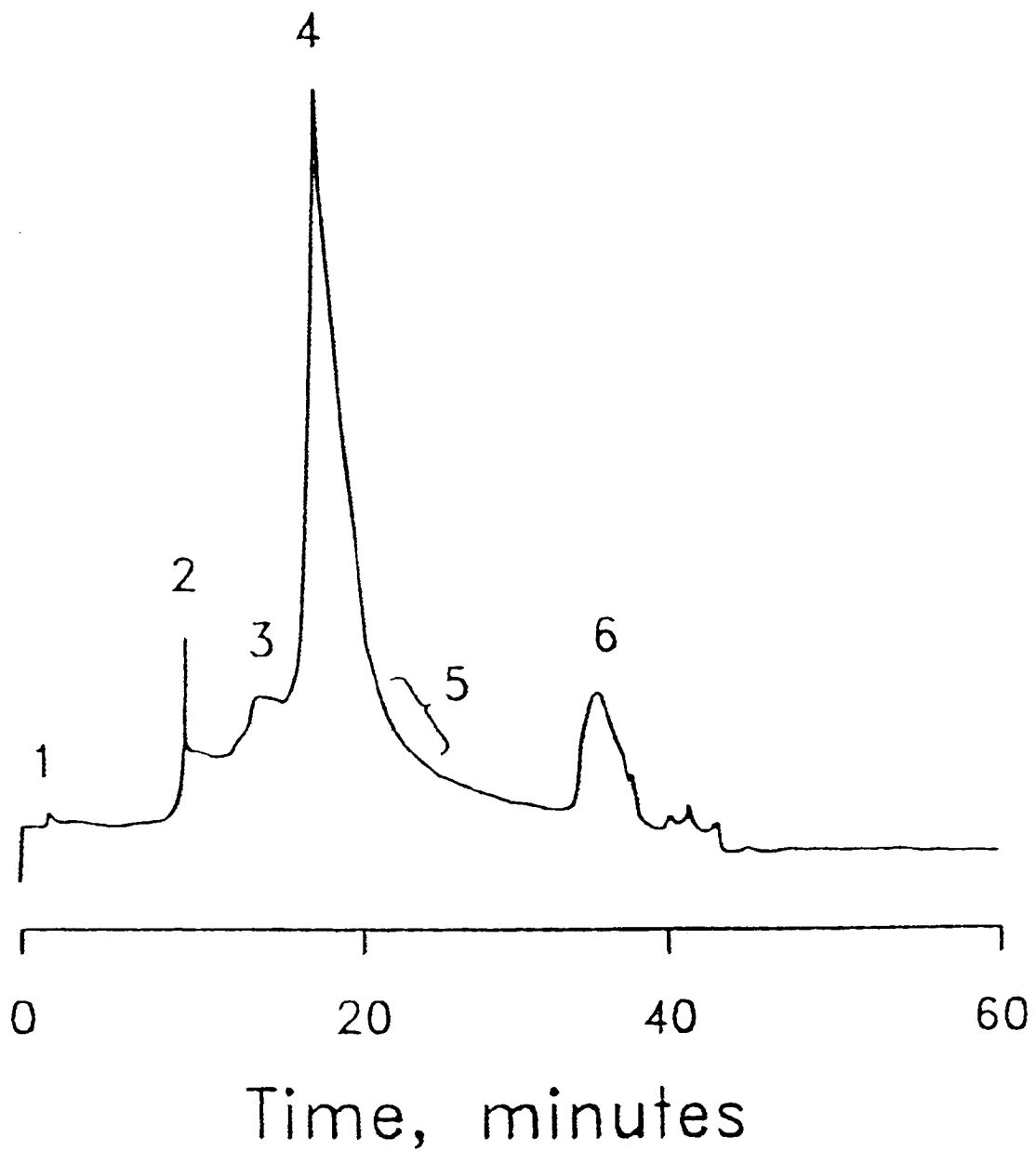


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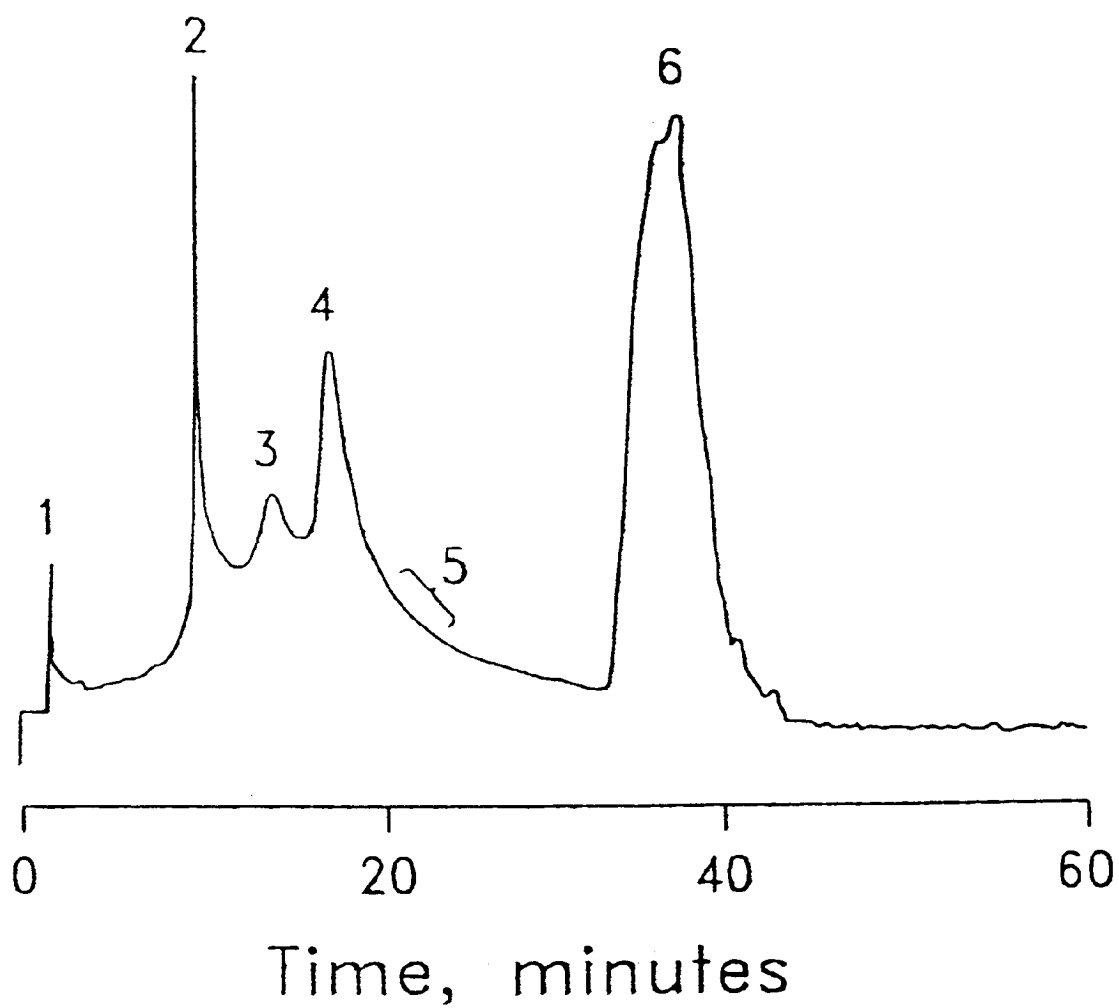
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*FIG. 1*



*FIG. 2*

## SYNTHETIC SOIL-EXTRACT MATERIALS AND MEDICAMENTS FOR INFLUENZA VIRUSES BASED THEREON

### CROSS-REFERENCE TO RELATED APPLICATION

This application claims the benefit under 35 U.S.C. §119 (e) of U.S. Provisional Application Ser. No. 60/297,669, filed Jun. 12, 2001 and is a continuation-in-part of application Ser. No. 09/345,865, filed Jul. 1, 1999, which is a divisional of application Ser. No. 08/798,329, filed Feb. 10, 1997, now U.S. Pat. No. 5,945,446, issued Aug. 31, 1999.

### BACKGROUND OF THE INVENTION

#### 1. Field of the Invention

This invention relates to synthetic soil extract substances comprised of phenolic polymers, and to compositions and methods for employing these synthetic phenolic polymers for preventing, reducing, treating, or eliminating influenza viral diseases.

#### 2. Description of Related Art

Soil extract materials, particularly the classes of substances known collectively as "humus," "humics," "humic acid(s)," or "humates," have been widely used in a number of applications for many years, as reviewed by F. J. Stevenson, *Humus Chemistry. Genesis Composition Reactions*; New York: Wiley, 1964; and, more recently, by A. Piccolo, *Humic Substances in Terrestrial Ecosystems*; New York: Elsevier, 1996.

Humic substances have long been known to exhibit anti-viral properties (H. Schultz, *Dtsch. Tierarztl. Wochenschr.* 1962, 69, 613; 1965, 72(13), 294-297; R. Klockung and M. Sprossig, *Experientia* 1972, 28(5), 607-608), particularly retroviruses (G. Sydow, V. Wunderlich, R. Klockung, and B. Helbig, *Pharmazie* 1986, 41(12), 865-868). Viral pathogens for which soil-extract materials have been shown to be effective include in particular Coxsackie virus A9 (Griggs-Baylor) (R. Klockung and M. Sprossig, *Experientia* 1972, 28(5), 607-608), herpes simplex virus type 1 (B. T. Rouse (Ed.), *Herpes Simplex Virus*; Berlin: Springer-Verlag, 1992; R. Klockung, K. D. Thiel, P. Wutzler, B. Helbig, and P. Drabke, *Pharmazie* 1978, 33(8), 539; F. Schiller, R. Klockung, P. Wutzler, and I. Farber, *Dermatol. Monatsschr.* 1979, 165(7), 505-509; B. Helbig, A. Sauerbrei, R. Klockung, P. Wutzler, N. Wicht, U. Wiedemann, and G. Herrmann, *J. Med. Virol.* 1987, 23(3), 303-309; R. Klockung and B. Helbig, in *Humic Substances in the Aquatic and Terrestrial Environment*; Berlin: Springer-Verlag, 1991; 407-412;) and type 2 (anon. *Zentralbl. Bakteriol [Orig. A]* 1976, 234(2), 159-169; K. D. Thiel, R. Klockung, H. Schweizer, and M. Sprossig, *Zentralbl. Bakteriol [Orig. A]* 1977, 239(3), 304-321; K. D. Thiel, B. Helbig, R. Klockung, P. Wutzler, M. Sprossig, and H. Schweizer, *Pharmazie* 1981, 36(1), 50-53; K. D. Thiel, B. Helbig, M. Sprossig, R. Klockung, and P. Wutzler, *Acta Virol.* 1983, 27(3), 200-208; K. D. Thiel, P. Wutzler, B. Helbig, R. Klockung, M. Sprossig, and H. Schweizer, *Pharmazie* 1984, 39(11), 781-782); human immunodeficiency virus (HIV) (M. Cushman, P. Wang, S. H. Chang, C. Wild, E. De Clercq, D. Schols, M. E. Goldman, and J. A. Bowen, *J. Med. Chem.* 1991, 34(1), 329-337; M. Cushman, S. Kanamathareddy, E. De Clercq, D. Schols, M. E. Goldman, and J. A. Bowen, *J. Med. Chem.* 1991, 34(1), 337-342; D. Schols, P. Wutzler, R. Klockung, B. Helbig, and E. De Clercq, *J. Acquir. Immune Defic. Syndr.* 1991, 4(7), 677-685; S. Loya, R. Tal, A. Hizi, S. Issacs, Y. Kashman, and Y. Loya, *J. Nat. Prod.* 1993, 56(12),

2120-2125; J. Schneider, R. Weis, C. Manner, B. Kary, A. Werner, B. J. Seubert, and U. N. Riede, *Virology* 1996, 218(2), 389-395; influenza virus type A (Krasnodar/101/59/H2N2) (R. Mentel, B. Helbig, R. Klockung, L. Dohner, and M. Sprossig, *Biomed. Biochim. Acta* 1983, 42(10), 1353-1356); and type B (J. Hils, A. May, M. Sperber, R. Klockung, B. Helbig, and M. Sprossig, *Biomed. Biochim. Acta* 1986, 45(9), 1173-1179); as well as other respiratory tract infectious agents (A. Jankowski, B. Nienartowicz, B. Polanska, and A. Lewandowicz-Uszyuska, *Arch. Immunol. Ther. Exp. (Warsz)* 1993, 41(1), 95-97).

The mechanisms whereby humic substances inhibit the cytopathicity of a number of viruses have been studied in some detail. It is thought that the materials prevent viral replication in part by sorbing onto the viral envelope protein (gp120 in the case of HIV), and thereby block the sorption of viral particles to cell surfaces: K. D. Thiel, R. Klockung, H. Schweizer, and M. Sprossig, *Zentralbl. Bakteriol. [Orig. A]* 1977, 239(3), 304-321; D. Schols, P. Wutzler, R. Klockung, B. Helbig, and E. De Clercq, *J. Acquir. Immune Defic. Syndr.* 1991, 4(7), 677-685; anon., *Fortschr. Med.* 1995, 113(7), 10; J. Schneider, R. Weis, C. Manner, B. Kary, A. Werner, B. J. Seubert, and U. N. Riede, *Virology* 1996, 218(2), 389-395. [Extracellular interception of pathogens by chemical agents that bind to them is a well-known means of immunological defense (D. M. Shankel, S. Kuo, C. Haines, and L. A. Mitscher, in *Antimutagenesis and Anticarcinogenesis Mechanisms III*; G. Bronzetti, H. Hayatsu, S. De Flora, M. D. Waters, and D. M. Shankel (Eds.); New York: Plenum, 1993; 65-74). Such materials might well be termed "despathogens", following the terminology proposed by T. Kada and K. Shimoi, *Bioessays* 1987, 7, 113-116, regarding "desmutagens".] It has also been found that naturally-occurring humic acid preparations can stimulate the production of cytokines, including interferon-gamma, interferon-alpha, and tumor necrosis factor-alpha (A. D. Inglot, J. Zielinska-Jencylyk, and E. Piasecki, *Arch. Immunol. Ther. Exp. (Warsz)* 1993, 41(1), 73-80); as well as interferon-beta (Z. Blach-Olszewska, E. Zaczynska, E. Broniarek, and A. D. Inglot, *Arch. Immunol. Ther. Exp. (Warsz)*, 1993, 41(1), 81-85).

The toxicity of naturally-occurring humic acids is remarkably low (K. D. Thiel, B. Helbig, R. Klockung, P. Wutzler, M. Sprossig, and H. Schweizer, *Pharmazie* 1981, 36(1), 50-53; U. N. Riede, I. Jonas, B. Kim, U. H. Usener, W. Kreutz, and W. Schlickewey, *Arch. Orthop. Trauma Surg.* 1992, 111(5), 259-264; H. Czyzewska-Szafran, Z. Jastrzebski, D. Soltysiak-Pawluczak, M. Wutkiewicz, A. Jedrych, and M. Riemiszewska, *Acta Pol. Pharm.* 1993, 50(4-5), 373-377; H. L. Yang, F. J. Lu, S. L. Wung, and H. C. Chiu, *Thromb. Haemost.* 1994, 71(3), 325-330). [Cytotoxic effects of anti-viral substances, including humic acids, are usually evaluated via biological (viability and alterations of cell morphology) and biochemical testing methods (<sup>51</sup>Cr release), as described by K. D. Thiel, U. Eichhom, H. Schweizer, and R. Klockung, *Arch. Toxicol. Suppl.* 1980, 4, 428-430.] The cytotoxicity (CD<sub>50</sub>) of a naturally-occurring humic acid for human peripheral blood leukocytes (PBL) was found to be 1-9 milligrams per milliliter. In addition, J. Schneider, R. Weis, C. Manner, B. Kary, A. Werner, B. J. Seubert, and U. N. Riede, *Virology* 1996, 218(2), 389-395, reported that the cytotoxicity of a synthetic humic acid prepared from hydroquinone for MT-2 cells was approximately 600 micrograms per milliliter. It has also been found that medicaments prepared from humic acids isolated from naturally-occurring soil materials are neither carcinogenic (Syrian hamster embryo cell transfor-

mation test: J. Koziorowska and E. Anuszevska, *Acta Pol. Pharm.* 1994, 51(1), 101-102) nor mutagenic (T. Sato, Y. Ose, and H. Hagase, *Mutat. Res.* 1986, 162(2), 173-178; V. M. Sui, A. I. Kiung, and T. I. Veidebaum, *Vopr. Kurortol. Fiozioter. Lech. Fiz. Kult.* 1986, 2(3-4), 34-37; J. Koziorowska, B. Chlopkiwicz, and E. Anuszevska, *Acta Pol. Pharm.* 1993, 50(4-5), 379-382). Prenatal (S. Golbs, V. Fuchs, M. Kuhnert, and C. Polo, *Arch. Exp. Veterinarmed.* 1982, 36(2), 179-185) and embryotoxic and teratogenic effects (T. Juskiewicz, M. Minta, B. Wlodarczyk, B. Biernacki, and J. Zmudzki, *Acta Pol. Pharm.* 1993, 50(4-5), 383-388) are also not observed with humic preparations at daily dose levels from 5-50 milligrams per kilogram body weight. Topical preparations are tolerated even better (V. V. Soldatov and M. N. Cherepanova, *Vopr. Kurortol. Fiozioter. Lech. Fiz. Kult.* 1970, 35(3), 256-259; H. Czyzewska-Szafran, Z. Jastrzebski, D. Soltysiak-Pawluczuk, M. Wutkiewicz, A. Jedrych, and M. Remiszewska, *Acta Pol. Pharm.* 1993, 50(4-5), 373-377) when applied dermally in aqueous solution in amounts as high as 10 percent weight-by-volume (K. Wiegleb, N. Lange, and M. Kuhnert, *Dtsch. Tierarztl. Wochenschr.* 1993, 100(10), 412-416).

Because humic substances are not chemically well-defined, the preparation of synthetic humic acids whose physicochemical properties mimic naturally-occurring materials is quite difficult, as pointed out by K. Murray and P. W. Linder, *J. Soil Sci.* 1983, 34, 511-523. Nevertheless, there have been several notable advances in this area. Broadly speaking, three general strategies have evolved. All depend upon starting with well-defined molecules of molecular weight on the order of hydroxybenzoic acid, and then causing the molecules to polymerize upon themselves to form larger molecules. The methods differ in the causation factor, which can be microbial, chemical, or enzymatic.

Humic acids of microbial origin have been described and discussed by M. Robert-Gero, C. Hardisson, L. Le Borgne, and G. Pignaud, *Ann. Inst. Pasteur (Paris)* 1966, 111(6), 750-767; and by M. Robert-Gero, C. Hardisson, L. Le Borgne, and G. Vidal, *Ann. Inst. Pasteur (Paris)* 1967, 113(6), 903-909.

The chemical synthesis of humic acids has been pioneered by R. Klocking, B. Helbig, and associates: R. Klocking, B. Helbig, and P. Drabke, *Pharmazie* 1977, 32, 297; R. Klocking, B. Helbig, K. D. Thiel, T. Blumohr, P. Wutzler, M. Sprossig, and F. Schiller, *Pharmazie* 1979, 34(5-6), 293-294; R. Mentel, B. Helbig, R. Klocking, L. Dohner and M. Sprossig, *Biomed. Biochim. Acta* 1983, 42(10), 1353-1356; H. P. Klocking, R. Klocking, and B. Helbig, *Farmakol. Toksikol.* 1984, 47(1), 93-95; K. D. Thiel, P. Wutzler, B. Helbig, R. Klocking, M. Sprossig, and H. Schweizer, *Pharmazie* 1984, 39(11), 781-782; J. Hils, A. May, M. Sperber, R. Klocking, B. Helbig, and M. Sprossig, *Biomed. Biochim. Acta* 1986, 45(9), 1173-1179; B. Helbig, A. Sauerbrei, R. Klocking, P. Wutzler, N. Wicht, U. Wiedemann, and G. Herrmann, *J. Med. Virol.* 1987, 23(3), 303-309; K. I. Hanninen, R. Klocking, and B. Helbig, *Sci. Total Environ.* 1987, 62, 201-210; R. Klocking and B. Helbig, in *Humic Substances in the Aquatic and Terrestrial Environment*; New York: Springer-Verlag, 1989; 407-412; C. Schewe, R. Klocking, B. Helbig, and T. Schewe, *Biomed. Biochim. Acta* 1991, 50(3), 299-305; D. Schols, P. Wutzler, R. Klocking, B. Helbig, and E. De Clercq, *J. Acquir. Immune Defic. Syndr.* 1991, 4(7), 677-685. Typically, 10 millimoles of the starting small-molecule phenolic compound is dissolved in distilled water, the pH is adjusted to 8.5 with aqueous sodium hydroxide (NaOH), and then 2-5 millimoles of sodium periodate (NaIO<sub>4</sub>) is added. The solution is

warmed at 50° C. for 30 minutes, and is then allowed to stand overnight. The resultant humic acid-like polymeric products are isolated by precipitation with lead(II) nitrate [Pb(NO<sub>3</sub>)<sub>2</sub>]. The precipitated polymers are redissolved in aqueous sodium hydroxide (pH 8.5) and heated with 8-hydroxyquinoline for 30 minutes at 100° C. The precipitate formed is lead(II) chelate, which is removed by filtration. Residual 8-hydroxyquinoline is extracted with chloroform, and the desired polymeric material is then precipitated from the aqueous solution by the addition of various combinations of acetic acid, ethyl acetate, and ethanol. Starting compounds that have been used for the synthesis of humic-like materials include 4-[bis(p-hydroxyphenyl)methylene]-2,5-cyclohexadien-1-one (aurin), 4-[bis(3-carboxy-4-hydroxyphenyl)methylene]-2-carboxy-2,5-cyclohexa-dien-2-one (aurintricarboxylic acid), 3-(3,4-dihydroxyphenyl)propenoic acid (caffeic acid), 1,2-dihydroxybenzene (catechol), 1,3,4,5-tetrahydroxycyclohexanecarboxylic acid 3-(3,4-dihydroxyphenyl)propenoate (chlorogenic acid), 3,4-dihydroxyphenylacetic acid (homoprotocatechuic acid), 1-(3,4-dihydroxyphenyl)-2-(N-methylamino)ethanol (epinephrine), 3-(4-hydroxy-3-methoxyphenyl)-2-propenoic acid (ferulic acid), 3,4,5-trihydroxybenzoic acid (gallic acid), 2,5-dihydroxybenzoic acid (gentic acid), 2,5-dihydroxyphenylacetic acid (homogentic acid), 3-(3,4-dihydroxyphenyl)propionic acid (hydrocaffeic acid), 1,4-dihydroxybenzene (hydroquinone), 2,3-dihydroxytoluene (3-methylcatechol), 3,4-dihydroxytoluene (4-methylcatechol), 2,5-dihydroxytoluene (2-methylhydroquinone), 4,4'-(2,3-dimethyltetramethylene)-di-(1,2-dihydroxybenzene) (nordihydroguaiaretic acid), 1-(3,4-dihydroxyphenyl)-2-aminoethanol (norepinephrine), 3,4-dihydroxybenzoic acid (protocatechuic acid), 1,2,3-trihydroxybenzene (pyrogallol), 1,3-dihydroxybenzene (resorcinol), and 4-hydroxy-3-methoxybenzoic acid (vanillic acid). Other notable efforts on the chemical synthesis of humic-like substances include the studies by De Clercq and colleagues on aurintricarboxylic acid, its derivatives, and related compounds: M. Cushman, P. Wang, S. H. Chang, C. Wild, E. De Clercq, D. Schols, M. E. Goldman, and J. A. Bowen, *J. Med. Chem.* 1991, 34(1), 329-337; M. Cushman, S. Kanamathareddy, E. De Clercq, D. Schols, M. E. Goldman, and J. A. Bowen, *J. Med. Chem.* 1991, 34(1), 337-342. Related efforts have also been reported by M. Robert-Gero, C. Hardisson, L. Le Borgne, and G. Vidal, *Ann. Inst. Pasteur (Paris)* 1967, 113(6), 903-909; M. Jakubiec, E. Miszczyk, and J. Szczerkowska, *Acta Microbiol. Pol. [B]* 1971, 3(1), 63-66; R. Ansorg and W. Rochus, *Arzneimittelforschung* 1978, 28(12), 2195-2198; J. Pommery, M. Imbenotte, A. F. Urien, D. Marzin, and F. Erb, *Mutat. Res.* 1989, 223(2), 183-189; F. J. Lu and Y. S. Lee, *Sci. Total Environ.* 1992, 114, 135-139; K. Wiegleb, N. Lange, and M. Kuhnert, *DTW Dtsch. Tierarztl. Wochenschr.* 1993, 100(10), 412-416; H. L. Yang, F. J. Lu, S. L. Wung, and H. C. Chiu, *Thromb. Haemost.* 1994, 71(3), 325-330; W. Seffner, F. Schiller, R. Heinze, and R. Breng, *Exp. Toxicol. Pathol.* 1995, 47(1), 63-70; and J. Schneider, R. Weis, C. Manner, B. Kary, A. Werner, B. J. Seubert, and U. N. Riede, *Virology* 1996, 218(2), 389-395.

The enzymatic catalytic synthesis of humic acids dates to about 1961 with the work by R. E. Hampton and R. W. Fulton, *Virology* 1961, 13, 44-52 (see also R. E. Hampton, *Phytopathology* 1970, 60, 1677-1681), who found that enzymatically oxidized phenols inactivate phytopathogenic (i.e., plant-related) viruses. Typically o-diphenol oxidase has



been employed for the enzymatic synthesis of humic-like materials: anon. *Zentralbl. Bakteriol. [Orig A]* 1976, 234(2), 159-169; R. Klocking, B. Helbig, and P. Drabke, *Pharmazie* 1977, 32(5), 297; K. D. Thiel, B. Helbig, R. Klocking, P. Wutzler, M. Sprossig, and H. Schweizer, *Pharmazie* 1981, 36(1), 50-53; K. D. Thiel, B. Helbig, M. Sprossig, R. Klocking, and P. Wutzler, *Acta Virol.* 1983, 27(3), 200-208; K. D. Thiel, P. Wutzler, B. Helbig, R. Klocking, M. Sprossig, and H. Schweizer, *Pharmazie* 1984, 39(11), 781-782; and G. Sydow, V. Wunderlich, R. Klocking, and B. Helbig, *Pharmazie* 1986, 41(12), 865-868.

A direct comparison of humic acids synthesized enzymatically and nonenzymatically from caffeic and hydrocaffeic acids has shown that the two synthetic routes produce materials that differ somewhat in their efficacy for the suppression of herpes (hominis) types 1 and 2 viruses: K. D. Thiel, P. Wutzler, B. Helbig, R. Klocking, M. Sprossig, and H. Schweizer, *Pharmazie* 1984, 39(11), 781-782.

PCT application WO 00/16785 (Mar. 20, 2000) from Dekker and Medlen discloses the use of humic acid or its salts, esters, or derivatives thereof, all prepared as described in U.S. Pat. Nos. 4,912,256 and 5,004,831 from coal extracts, in stimulating lymphocytes in a human, animal, or bird. This allows for the treatment of viral and bacterial infections, and more particularly HIV infections, cancer, and opportunistic diseases. Oxihumic acids, salts, esters, or derivatives thereof are preferred. Administration is preferably oral. Some example pharmacological data presented include the antiviral activity of oxihumates against HIV-1 in vitro and clinical trials of oral oxihumate in HIV-infected patients.

PCT application WO 00/16786 (Mar. 30, 2000) from Dekker and Medlen discloses the use of pharmaceutical compositions comprising an oxihumic acid or its salts, esters, or derivatives thereof, all prepared as described in U.S. Pat. Nos. 4,912,256 and 5,004,831 from coal extracts, as active ingredients. Compositions are preferably administered orally for stimulating lymphocytes in a human, animal, or bird. They may be used in treating viral and bacterial infections, HIV infections, opportunistic diseases, inflammation, pain and fever, cancer growth, and diseases associated with viral infection and a depressed immune system. A number of pharmacological examples are given, including interleukin 10 production by oxihumate-treated lymphocytes, increased antibody production against Newcastle disease in chickens treated with oxihumate, TNF production by oxihumate-treated lymphocytes, and antiviral activity of oxihumate against HSV-1 and coxsackie virus type 1 in vitro.

The diversity of physicochemical characteristics as well as wide variation in the biological activity and toxicity of humics extracted or otherwise derived from natural soils has been well documented. This diversity and variation is due to variations in factors such as the source of the soil, the method(s) of extraction and/or isolation, and the technique(s) employed to treat the extract once it has been separated and isolated from crude soil. The consequence of irreproducibility of the properties of substances extracted from natural soil is that the commercial value of such materials is minimized. In addition, they are rendered unsuitable as medicaments. Also, while a number of laboratory-scale processes have already been described that address various aspects of the isolation, synthesis, and/or preparation of humic substances or similar materials, there are no reports of preparing and isolating such purely synthetic humic acids or similar materials by methods that are suitable for scaleup directly to industrial levels, that provide economically

acceptable yields, and that optimize the preparation procedures from the standpoint of medicament safety and efficacy. All of the known synthetic methods utilize potentially toxic precipitation methods [lead(II) nitrate precipitation] followed by complex isolation procedures, potentially mutagenic compound-producing hydrochloric acid precipitation or lengthy synthetic steps as long as 10 days.

#### SUMMARY OF THE INVENTION

There is a need to devise simple synthetic procedures that yield inexpensive, safe materials whose physicochemical attributes are reproducible, and that at least simulate those of typical commercially-available soil extracts.

One embodiment is a method for preventing and/or treating influenza virus infection in a mammal which comprises administering an effective amount of a synthetic phenolic polymeric material which is prepared by:

- a) dissolving in an aqueous solution at least one starting organic compound comprising at least one hydroxyl group and at least one carbonyl group or at least two hydroxyl groups on an aromatic structure;
- b) adjusting the pH of the aqueous solution resulting from step a) to between about 8 and 11;
- c) adding an alkaline periodate salt or alkaline-earth periodate salt to the aqueous solution resulting from step b);
- d) maintaining the temperature of the solution from step c) between about 20° C. and 100° C. for a period of at least about 30 minutes;
- e) adding at least one water soluble compound or salt selected from the group consisting of boric acid, borate salts, alkaline earth salts, transition metal salts, alkaline sulfides, alkaline earth sulfides, or transition metal sulfides to the aqueous solution resulting from step d);
- f) allowing the aqueous solution from step e) to stand with or without stirring at about 20° C. to 100° C. for at least about 2 hours; and
- g) removing molecules from the solution resulting from step f) below about 500 to about 10,000 daltons.

In another aspect, the method of preparation of the synthetic phenolic polymeric material further comprises a step, following the step of removing molecules from the solution below about 500 daltons to 10,000 daltons, of concentrating the solution.

In another aspect, the method of preparation of the synthetic phenolic polymeric material further comprises a step, following the step of removing molecules from the solution below about 500 dalton to 10,000 daltons, of removing water from the solution.

In another aspect, the influenza virus infection is effected by a virus, preferably influenza A or influenza B.

In another aspect, the administering of a synthetic phenolic polymeric material is along with an effective amount of an antiviral composition.

In another aspect, the administering of a synthetic phenolic polymeric material can be achieved systemically or topically.

One embodiment is a method for inhibiting influenza viral attachment to host cells in a mammal which comprises administering an effective amount of a synthetic phenolic polymeric material which is prepared by:

- a) dissolving in an aqueous solution at least one starting organic compound comprising at least one hydroxyl group and at least one carbonyl group or at least two hydroxyl groups on an aromatic structure;

7

- b) adjusting the pH of the aqueous solution resulting from step a) to between about 8 and 11;
- c) adding an alkaline periodate salt or alkaline-earth periodate salt to the aqueous solution resulting from step b);
- d) maintaining the temperature of the solution from step c) between about 20° C. and 100° C. for a period of at least about 30 minutes;
- e) adding at least one water soluble compound or salt selected from the group consisting of boric acid, borate salts, alkaline earth salts, transition metal salts, alkaline sulfides, alkaline earth sulfides, or transition metal sulfides to the aqueous solution resulting from step d);
- f) allowing the aqueous solution from step e) to stand with or without stirring at about 20° C. to 100° C. for at least about 2 hours; and
- g) removing molecules from the solution resulting from step f) below about 500 to about 10,000 daltons.

In another aspect, the method of preparation of the synthetic phenolic polymeric material further comprises a step, following the step of removing molecules from the solution below about 500 daltons to 10,000 daltons, of concentrating the solution.

In another aspect, the method of preparation of the synthetic phenolic polymeric material further comprises a step, following the step of removing molecules from the solution below about 500 dalton to 10,000 daltons, of removing water from the solution.

#### BRIEF DESCRIPTION OF THE DRAWINGS

FIG. 1 shows a high-performance liquid chromatography (HPLC) trace for the synthetic humic acid product obtained from 2,5-dihydroxyphenylacetic acid (homogentisic acid), as described in Examples 10, 11, and 14.

FIG. 2 shows a high-performance liquid chromatography (HPLC) trace obtained for a typical commercially-available natural-product humic acid.

#### DETAILED DESCRIPTION OF THE PREFERRED EMBODIMENTS

U.S. Pat. No. 5,945,446, issued Aug. 31, 1999, discloses the process for preparing synthetic soil-extract materials and medicaments based thereon. U.S. Pat. No. 5,945,446 and the references therein are incorporated herein by reference in their entirety.

The inventor has developed combinations of chemical processes for the preparation of synthetic phenolic polymeric materials, also known as synthetic humic acids, whose physicochemical properties and attributes are reproducible, and which simulate those of typical commercially-available natural humic acids and other soil extracts, which contain little or no ionic salts or other compounds of molecular weight less than about 500 daltons, which have a minimum molecular weight of about 500 daltons, and which processes shall be suitable for scaleup directly to industrial levels that provide economically acceptable yields.

The inventor has also developed compositions and methods for treating or preventing influenza viral diseases by using an effective amount for anti-viral activity of a synthetic humic acid prepared according to the above processes.

The starting compounds used in the chemical processes employed for production of synthetic humic acids according to embodiments of the method of the preferred embodiments are known materials that are readily available commercially.

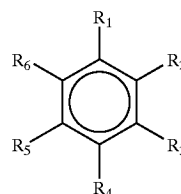
Generally speaking, the chemical processes for the preparation of synthetic humic acids of the preferred embodiments include the following steps:

8

- a) dissolving in an aqueous solution at least one starting organic compound comprising at least one hydroxyl group and at least one carbonyl group or at least two hydroxyl groups on an aromatic structure;
- b) adjusting the pH of the aqueous solution resulting from step a) to between about 8 and 11;
- c) adding an alkaline periodate salt or alkaline-earth periodate salt to the aqueous solution resulting from step b);
- d) maintaining the temperature of the solution from step c) between about 20° C. and 100° C. for a period of at least about 30 minutes;
- e) adding at least one water soluble compound or salt selected from the group consisting of boric acid, borate salts, alkaline earth salts, transition metal salts, alkaline sulfides, alkaline earth sulfides, or transition metal sulfides to the aqueous solution resulting from step d);
- f) allowing the aqueous solution from step e) to stand with or without stirring at about 20° C. to 100° C. for at least about 2 hours; and
- g) removing molecules from the solution resulting from step f) below about 500 to about 10,000 daltons;
- h) concentrating the solution resulting from step g), if necessary; and
- i) removing the water from the solution resulting from step h), if necessary.

The starting organic compound in step a) above can be one, or more than one in combination, of different compounds taken from the group consisting of starting organic compounds illustrated in Tables 1 and 2. Starting organic compounds illustrated in Table 1 are comprised of a single benzene ring with six substituents R<sub>1</sub>-R<sub>6</sub>, wherein R<sub>1</sub>-R<sub>6</sub> can be any one of the indicated atom or functional groups, as long as at least one of R<sub>1</sub>-R<sub>6</sub> is a hydroxy (—OH) functional group. Preferably, at least one of R<sub>1</sub>-R<sub>6</sub> is a hydroxy (—OH) functional group and at least one of the remaining substituents R<sub>1</sub>-R<sub>6</sub> contains a carboxylic acid functional group. More preferably, two of R<sub>1</sub>-R<sub>6</sub> are hydroxy (—OH) functional groups and one of the remaining substituents R<sub>1</sub>-R<sub>6</sub> contains a carboxylic acid functional group. Homogentisic acid, which is 2,5-dihydroxyphenylacetic acid, is a particularly preferred starting organic compound. Caffeic acid, which is 3,4-dihydroxycinnamic acid, is another particularly preferred starting organic compound. Chlorogenic acid, which is 1,3,4,5-tetrahydroxycyclohexanecarboxylic acid 3-(3,4-dihydroxycinnamate) is yet another particularly preferred starting organic compound.

TABLE 1



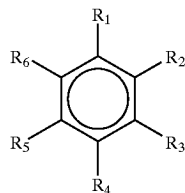
R<sub>1</sub>, R<sub>2</sub>, R<sub>3</sub>, R<sub>4</sub>, R<sub>5</sub>, R<sub>6</sub> =

—H  
—CH<sub>3</sub>  
—CH<sub>2</sub>CH<sub>3</sub>  
—(CH<sub>2</sub>)<sub>2</sub>CH<sub>3</sub>

9

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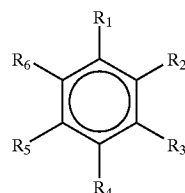
TABLE 1-continued



R<sub>1</sub>, R<sub>2</sub>, R<sub>3</sub>, R<sub>4</sub>, R<sub>5</sub>, R<sub>6</sub> =

- CH(CH<sub>3</sub>)<sub>2</sub>
- OH
- OCH<sub>3</sub>
- CHO
- CO<sub>2</sub>H
- CO<sub>2</sub>CH<sub>3</sub>
- CH<sub>2</sub>OH
- CH<sub>2</sub>OCH<sub>3</sub>
- CH<sub>2</sub>CHO
- CH<sub>2</sub>CO<sub>2</sub>H
- CH<sub>2</sub>CO<sub>2</sub>CH<sub>3</sub>
- (CH<sub>2</sub>)<sub>2</sub>OH
- (CH<sub>2</sub>)<sub>2</sub>OCH<sub>3</sub>
- (CH<sub>2</sub>)<sub>2</sub>CHO
- (CH<sub>2</sub>)<sub>2</sub>CO<sub>2</sub>H
- (CH<sub>2</sub>)<sub>2</sub>CO<sub>2</sub>CH<sub>3</sub>
- CH(CH<sub>3</sub>)OH
- CH(CH<sub>3</sub>)OCH<sub>3</sub>
- CH(CH<sub>3</sub>)CHO
- CH(CH<sub>3</sub>)CO<sub>2</sub>H
- CH(CH<sub>3</sub>)CO<sub>2</sub>CH<sub>3</sub>
- CH(CH<sub>3</sub>)CH<sub>2</sub>OH
- CH(CH<sub>3</sub>)CH<sub>2</sub>OCH<sub>3</sub>
- CH(CH<sub>3</sub>)CH<sub>2</sub>CHO
- CH(CH<sub>3</sub>)CH<sub>2</sub>CO<sub>2</sub>H
- CH(CH<sub>3</sub>)CH<sub>2</sub>CO<sub>2</sub>CH<sub>3</sub>
- CH(OH)<sub>2</sub>
- CH(OH)OCH<sub>3</sub>
- CH(OH)CHO
- CH(OH)CO<sub>2</sub>H
- CH(OH)CO<sub>2</sub>CH<sub>3</sub>
- CH(OCH<sub>3</sub>)OH

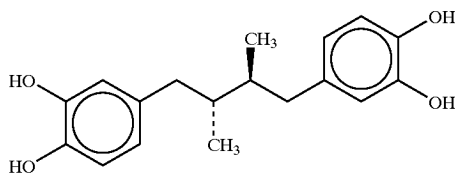
TABLE 1-continued



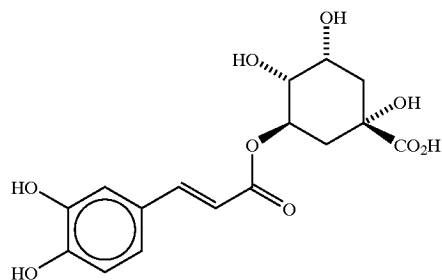
R<sub>1</sub>, R<sub>2</sub>, R<sub>3</sub>, R<sub>4</sub>, R<sub>5</sub>, R<sub>6</sub> =

- CH(OCH<sub>3</sub>)<sub>2</sub>
- CH(OCH<sub>3</sub>)CHO
- CH(OCH<sub>3</sub>)CO<sub>2</sub>H
- CH(OCH<sub>3</sub>)CO<sub>2</sub>CH<sub>3</sub>
- CH(OH)CH<sub>2</sub>OH
- CH(OH)CH<sub>2</sub>OCH<sub>3</sub>
- CH(OH)CH<sub>2</sub>CHO
- CH(OH)CH<sub>2</sub>CO<sub>2</sub>H
- CH(OH)CH<sub>2</sub>CO<sub>2</sub>CH<sub>3</sub>
- CH(OCH<sub>3</sub>)CH<sub>2</sub>OH
- CH(OCH<sub>3</sub>)CH<sub>2</sub>OCH<sub>3</sub>
- CH(OCH<sub>3</sub>)CH<sub>2</sub>CHO
- CH(OCH<sub>3</sub>)CH<sub>2</sub>CO<sub>2</sub>H
- CH(OCH<sub>3</sub>)CH<sub>2</sub>CO<sub>2</sub>CH<sub>3</sub>
- (CH<sub>2</sub>)<sub>3</sub>OH
- (CH<sub>2</sub>)<sub>3</sub>OCH<sub>3</sub>
- (CH<sub>2</sub>)<sub>3</sub>CHO
- (CH<sub>2</sub>)<sub>3</sub>CO<sub>2</sub>H
- (CH<sub>2</sub>)<sub>3</sub>CO<sub>2</sub>CH<sub>3</sub>
- CHCHOH (cis or trans)
- CHCHOCH<sub>3</sub> (cis or trans)
- CHCHCHO (cis or trans)
- CHCHCO<sub>2</sub>H (cis or trans)
- CHCHCO<sub>2</sub>CH<sub>3</sub> (cis or trans)
- CH<sub>2</sub>CHCHOH (cis or trans)
- CH<sub>2</sub>CHCHOCH<sub>3</sub> (cis or trans)
- CH<sub>2</sub>CHCHCHO (cis or trans)
- CH<sub>2</sub>CHCHCO<sub>2</sub>H (cis or trans)
- CH<sub>2</sub>CHCHCO<sub>2</sub>CH<sub>3</sub> (cis or trans)

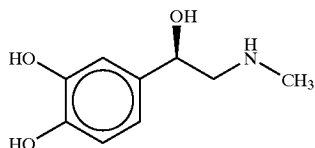
TABLE 2



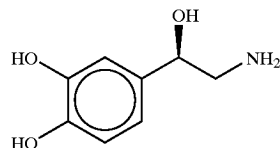
Nordihydroguaiaretic Acid



Chlorogenic Acid

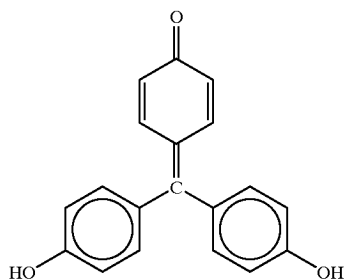


Epinephrine

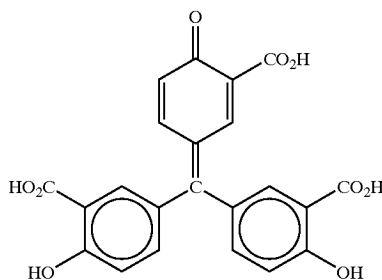


Norepinephrine

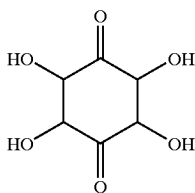
TABLE 2-continued



Aurin



Aurintricarboxylic Acid



Tetrahydroxybenzoquinone

Various initial concentrations of starting organic compounds in water, preferably distilled, can be employed and no lower or upper limits are uniformly required. A low concentration solution of sodium hydroxide, such as about 0.1 Normal, may also be employed as a diluent for the starting organic compound. The appropriate initial concentration of the starting organic compound or compounds is determined by the synthesis yield requirements and inherent requirements, such as the upper limit of aqueous solubility of the starting organic compound or compounds. Conventional methods are employed to determine the appropriate initial concentration of the starting organic compound or compounds.

The pH of the aqueous solution containing the starting organic compound or compounds can be adjusted in step b) to between about 8 and 11 by adding aqueous ammonium hydroxide, or other aqueous alkaline oxide or hydroxide, or aqueous alkaline earth oxide or hydroxide, or aqueous transition metal oxide or hydroxide. Additionally, if the initial aqueous solution contains a low concentration of base, such as about 0.1 Normal sodium hydroxide and the initial solution pH is too high, an acid such as hydrochloric acid may be employed to adjust the pH to the desired value. Other inorganic acids may also be employed for pH adjustment. Note that if hydrochloric acid is employed to adjust the pH downwards from an initial high value, care should be taken to avoid letting the pH go below about 8. Acidic conditions below about pH 7 should be avoided in the presence of hydrochloric acid to eliminate the possibility of formation of mutagenic chlorinated humic acid materials.

An alkaline periodate salt or alkaline earth periodate salt may be employed as an oxidant or polymerization initiator of the starting organic compound in step c). Sodium periodate is particularly preferred. The concentration of the alkaline periodate salt or alkaline earth periodate salt is generally between about 10% and 100% of the starting organic compound or compounds on a molar basis. Thus, if 10 millimoles of starting organic compound is employed, 1 to 10 millimoles of alkaline periodate salt may be employed. Preferably, a molar concentration of periodate that is about 10%–50% of the molar concentration of the starting organic

compound or compounds is employed. Most preferably, a molar concentration of periodate that is about 25%–35% of the molar concentration of the starting organic compound or compounds is employed. The exact concentration to be used can be determined by conventional synthetic yield optimization techniques.

Alkaline or alkaline earth sulfides or transition metal sulfides can be optionally added to the initial aqueous solution containing the starting organic compound or compounds following the pH adjustment in step b) and immediately before, at the same time, or following the addition of the periodate in step c). Sulfides contribute to the phenolic polymeric structure, the stability of the structure and its biological activity. Sodium sulfide nonahydrate is a particularly preferred sulfide. The concentration of the sulfide is generally between about 1% and 20% of the starting organic compound or compounds on a molar basis. Thus, if 10 millimoles of starting organic compound is employed, 0.1 to 2 millimoles of sulfide may be employed. Preferably, a molar concentration of sulfide that is about 5%–15% of the molar concentration of the starting organic compound or compounds is employed. Most preferably, a molar concentration of sulfide that is about 8% to 12% of the molar concentration of the starting organic compound or compounds is employed. The exact concentration of sulfide to be used can be determined by conventional synthetic yield optimization techniques.

Steps b) and c) above give conditions for oxidizing and polymerizing the starting organic compound. Although the use of periodate salt in basic conditions is preferable, there are other conditions that can perform oxidation and polymerization of the starting organic compound. One may substitute other reagents known in the art that are known to perform this function. If the reagents for oxidation and polymerization are substituted, the temperature and time period for the reaction in step d) should be adjusted accordingly for optimization. For example, a phenolic solution with about 2 equivalents of hydrogen peroxide can react for about one week at about 23° C. to form humic acids.

The pH-adjusted aqueous solution containing the starting organic compound, periodate and optional sulfide is placed

in a water-bath or other thermostat heating device at about 20° C. to 100° C. for a period of about 30 minutes to 100 hours in step d). Alternatively, the aqueous solution itself may be thermostated between about 20° C. and 100° C. for a period of about 30 minutes to 100 hours. A preferred temperature and time period is between about 35° C. and 80° C. for about 30 minutes to 100 hours. A particularly preferred temperature and time is about 50° C. for about 30 minutes to two hours. Alternative temperatures and pressures that are equivalent to the above temperature and pressures may be used.

Following this period, salts are added to the solution resulting from step d) alone or in combination in step e). Salts containing boron, calcium and other alkaline earths, iron and other transition metals are preferred. Such salts may contribute to the phenolic polymeric structure, its stability and biological activity. Boric acid or boron-containing-borate salts, such as sodium borate, are particularly preferred, as are alkaline earth salts, such as calcium sulfate dihydrate, and transition metal salts, such as ferrous sulfate heptahydrate. The concentrations of each of the salts employed is generally between about 0.1% and 20% of the starting organic compound or compounds on a molar basis. Preferably, a molar concentration of salt which is about 0.2% to 10% of the molar concentration of the starting organic compound or compounds is employed. Most preferably, a molar concentration of salt that is about 0.2% to 2% of the molar concentration of the starting organic compound or compounds, is employed. The exact concentration to be used can be determined by conventional synthetic yield optimization techniques. The solution resulting from step e) is allowed to stand at between about 20° C. and 100° C. with or without stirring for a period of at least 2 hours in step f). Preferably, the solution is allowed to stand at between about 20° C. and 80° C. for about 2 to 48 hours. Alternative temperatures and pressures that are equivalent to the above temperature and pressures may be used. Any precipitate formed at this stage is removed via conventional centrifugation.

Molecules below about 500 to about 10,000 daltons in the solution resulting from step f) are removed in step g). A variety of known conventional techniques can be employed, such as preparative chromatography, ultrafiltration or dialysis. Molecules are preferably removed from the solution resulting from step f) by employing dialysis in step g) with a flow-through open-channel or screen membrane apparatus consisting of a sandwich-type membrane of lower molecular-weight cutoff of 500–10,000 daltons until the conductivity of the solution has dropped to about 200 microsiemens or less. Most preferably, molecules are removed from the solution resulting from step f) by employing dialysis in step g) until the conductivity of the solution has dropped to about 50 microsiemens or less. A Pall Filtron Ultrasette® Tangential Flow Device or Mini-Ultrasette® Tangential Flow Device used with a Pall Filtron Ultralab® Specialized Pump and Reservoir System are preferred for solution dialysis.

The conductivity of the solution processed in step g) above can conveniently be monitored with a flow-through conductivity cell and conductivity meter. Alternatively, a simple inexpensive hand-held combination conductivity cell/conductivity meter (e.g., a Nalcometer Model MLN) can be employed.

Optionally, the solution from step g) above can be concentrated. One method of concentrating the solution is by removing water from the solution. The removal of water can be accomplished by conventional means known in the art.

Before removing the water from the solution in step h) above, the solution resulting from step g) above can be further dialyzed with a flow-through apparatus consisting of a sandwich-type membrane of molecular weight cutoff of 50,000 daltons. In this case the filtrate solution, not the retentate, is saved for further concentrating and processing according to steps h) and i). The resultant product will have a molecular-weight range of about 500–50,000 daltons.

If the solution resulting from either steps g) or h) above is to be stored as an aqueous solution for long periods of time for later application or use, for example as an anti-viral treatment solution, anti-viral therapy, anti-microbial therapy, a spray-on fertilizer or soil amendment, it can be filtered through standard 0.2–0.4 micron filters to remove bacteria and viruses, that is, can be made sterile by filtration. Alternatively, the aqueous solution from either steps g) or h) can be autoclaved for about 5–60 minutes at about 100–150° C. to produce a sterile solution.

A final optional step i) in the process of the preferred embodiments involves removing water from the solution resulting from step h). When freeze-drying is employed as the method of water removal in step i) above, the resultant product is a light fluffy dark-colored powder that is subject to static electricity effects. To minimize these effects, a small amount of mannose or other sugar can be added to the solution resulting from step h) just prior to freeze-drying. Water removal from the product can be carried out by means other than freeze-drying in step i) above, such as by heat evaporation with or without vacuum, by rotary evaporation, by spray-drying, or by any other solvent-removal technique that is convenient as well as economical for aqueous solutions. The dried powder obtained from step i) above can be autoclaved for about 15–30 minutes at about 100–120° C. to produce a sterile powder.

The synthetic humic acid materials produced according to the chemical processes and separation and isolation procedures of the preferred embodiments exhibit the physicochemical properties and attributes of typical naturally-occurring commercially-available humic acids and other soil extracts.

A facile method of examining the physicochemical characteristics of the product yielded by steps a) through h) above, or by modifications thereto, is high performance liquid chromatography (HPLC). The chromatographic fingerprint pattern so obtained from HPLC also offers a convenient means of comparing one product with another, as well as comparing each of the synthetic products with naturally-occurring humic acids and other soil-extract materials. The HPLC method is thus used to determine the reproducibility of the physicochemical properties and attributes of the synthetic phenolic polymeric materials, as well as to determine if the aforementioned properties and attributes simulate the physicochemical properties and attributes of typical commercial-available natural humic acids and other soil extracts. The latter determination of simulation is done in the conventional manner employing HPLC; e.g., by visually and quantitatively comparing the HPLC chromatographic fingerprint patterns of the materials. The fingerprint patterns of the two materials, one synthetic and one natural, need not be 100% identical to conclude that the physicochemical properties and attributes of the synthetic phenolic polymeric material simulates the physicochemical properties and attributes of the natural humic acid. An approximate correspondence between the aforementioned HPLC fingerprint patterns is all that is required to conclude that the synthetic material simulates the natural material. In general, even about 75% visual correspondence in two

HPLC fingerprint patterns is all that is necessary to conclude that one material simulates another.

A useful fingerprint pattern for natural as well as synthetic soil extract materials can be obtained as follows. The column is comprised of a packing, typically reversed-phase polymer PRP-1 (Hamilton Co.), of particle size 5 microns, and being 150 millimeters in length by 4.1 millimeters inside diameter. The mobile phase is comprised of three solutions. Solution A is about 0.1 Normal aqueous sodium hydroxide. Solution B is about 0.05 Normal of so-called Prideaux universal buffer, which is made by combining about 4.25 grams of sodium nitrate ( $\text{NaNO}_3$ ), about 12.37 grams of boric acid ( $\text{H}_3\text{BO}_3$ ), about 23.06 grams of phosphoric acid ( $\text{H}_3\text{PO}_4$ ), and about 12.01 grams of acetic acid ( $\text{CH}_3\text{CO}_2\text{H}$ ) with about 4 liters of distilled water. Solution C is about 100% methanol ( $\text{CH}_3\text{OH}$ ). The mobile-phase gradient employed for an HPLC run consists of about 40% solution A plus about 60% solution B at the beginning, which composition is changed in a linear manner to next about 100% solution A after about 20 minutes. The mobile phase is then changed linearly again to about 10% A plus about 90% C over the next about 5 minutes, which final composition is held for the purpose of a column wash for the next about 35 minutes. The mobile-phase flow rate is about 1 milliliter per minute. The detector is UV-Visible, which is set at 340 nanometers. The chart speed is typically 0.5 centimeter per minute. The sample loop size is about 5–20 microliters. Solutions are prepared for HPLC by dissolving about 0.1–10 grams of dried sample in about 100 milliliters of distilled water or about 0.1 Normal aqueous sodium hydroxide of about pH 8–10.

The chemical processes and separation and isolation procedures of the present invention are suitable for scaleup directly to industrial levels that provide economically acceptable yields. The chemical processes and separation and isolation procedures of the present invention can produce synthetic product yields approaching 100%. More typically, approximately 0.08 to 0.65 g of synthetic humic acid can be produced from about 10 millimoles of starting organic compound or compounds in about 300 milliliters. These procedures can be scaled up to pharmaceutical production scales employing about 10,000 to 20,000 liters or more of initial solution containing the starting organic compound or compounds. A total yield between approximately 2.7 and 21.7 kilograms of synthetic humic acid can be achieved utilizing a 10,000-liter thermally-jacketed stainless-steel tank and a concentration of starting organic compound of about 10 millimoles per about 300 milliliters. A single anti-viral treatment may employ milligram amounts of synthetic humic acid. Twenty kilograms of synthetic humic acid represents about 2 million units of anti-viral product at about 10 milligrams per unit. Even at a treatment cost of about \$0.10 per unit, this amount represents about \$200,000.00 of synthetic humic acid. Since the starting organic compounds utilized in the preferred embodiments are relatively inexpensive, the synthesis yields of the chemical processes and separation and isolation procedures of the preferred embodiments are economically very acceptable.

Examples 1 through 9 are illustrative of the variety of starting organic compounds that can be employed in the process of the preferred embodiments. It was not considered necessary to carry out all steps of the process of the preferred embodiments to illustrate starting compound variety. More particularly, Examples 1 through 9 are illustrative of all steps of the process of the preferred embodiments with the exception of step e), the addition of salts.

#### EXAMPLE 1

##### Preparation of a Synthetic Humic Acid from 2,5-dihydroxybenzoic Acid (Gentisic Acid)

The starting organic compound is shown in Table 1, and represented by  $\text{R}_1=\text{CO}_2\text{H}$ ;  $\text{R}_2$ ,  $\text{R}_5=\text{OH}$ ; and  $\text{R}_3$ ,  $\text{R}_4$ ,

$\text{R}_6=\text{H}$ . Ten millimoles (1.55 grams) of gentisic acid was dissolved in 300 milliliters of 0.1 Normal aqueous sodium hydroxide (NaOH). The solution pH was adjusted to 8.5 with 6 Normal HCl. Two and one-half millimoles (0.54 gram) of sodium periodate ( $\text{NaIO}_4$ ) was added, and the solution was placed in a water-bath at 50° C. for 30 minutes. The solution was allowed to stand at room temperature overnight. Any precipitate was removed by centrifugation. The solution was dialyzed with a 1,000-dalton cut-off flow-through open-channel or screen membrane system (Pall Filtron: Ultrasette® 7 Tangential Flow Device or Mini-Ultrasette® 7 Tangential Flow Device used with a Pall Filtron Ultralab® 7 Specialized Pump and Reservoir System) to a conductivity of 30 microsiemens or less against distilled water. The dialysis apparatus was then used to concentrate the solution to about 200 milliliters. The solution can be saved at this point for further use as an aqueous solution; or it can be freeze-dried to a powder. (Five to twenty hundredths of a gram of mannose or other suitable carbohydrate can be added to the solution prior to freeze-drying to reduce static electricity effects associated with the freeze-dried powder.) The yield of synthetic soil extract was 0.2 gram.

The following Examples 2–9 employ the synthesis procedure of Example 1 beginning with the adjustment of solution pH.

#### EXAMPLE 2

##### Preparation of a Synthetic Humic Acid from 3,4-dihydroxyphenylacetic Acid (Homoprotocatechuic Acid)

The starting organic compound, 3,4-dihydroxyphenylacetic acid, is shown in Table 1, and represented by  $\text{R}_1=\text{CH}_2\text{CO}_2\text{H}$ ;  $\text{R}_3$ ,  $\text{R}_4=\text{OH}$ ; and  $\text{R}_2$ ,  $\text{R}_5$ ,  $\text{R}_6=\text{H}$ . Ten millimoles (1.68 grams) of homoprotocatechuic acid was dissolved in 300 milliliters of 0.1 Normal aqueous sodium hydroxide (NaOH). The remaining procedure followed that of Example 1. The yield of synthetic soil extract was 0.24 gram.

#### EXAMPLE 3

##### Preparation of a Synthetic Humic Acid from dl-(3,4-dihydroxyphenyl)hydroxyacetic Acid (dl-3,4-dihydroxymandelic Acid)

The starting organic compound, dl-(3,4-dihydroxyphenyl)hydroxyacetic acid, is shown in Table 1, and represented by  $\text{R}_1=\text{CH}(\text{OH})\text{CO}_2\text{H}$ ;  $\text{R}_3$ ,  $\text{R}_4=\text{OH}$ ; and  $\text{R}_2$ ,  $\text{R}_5$ ,  $\text{R}_6=\text{H}$ . Ten millimoles (1.84 grams) of dl-3,4-dihydroxymandelic acid was dissolved in 300 milliliters of 0.1 Normal aqueous sodium hydroxide (NaOH). The remaining procedure followed that of Example 1. The yield of synthetic soil extract was 0.08 gram.

#### EXAMPLE 4

##### Preparation of a Synthetic Humic Acid from Aurintricarboxylic Acid

The chemical structure of the starting organic compound is shown in Table 2. Ten millimoles (4.2 grams) of aurintricarboxylic acid was dissolved in 300 milliliters of 0.1 Normal aqueous sodium hydroxide (NaOH). The remaining procedure followed that of Example 1. The yield of synthetic soil extract was 4.7 grams.

#### EXAMPLE 5

##### Preparation of a Synthetic Humic Acid from 3-(3,4-dihydroxyphenyl)propenoic Acid (Caffeic Acid)

The starting organic compound is shown in Table 1, and represented by  $\text{R}_1=\text{CHCHCO}_2\text{H}$ ;  $\text{R}_3$ ,  $\text{R}_4=\text{OH}$ ; and  $\text{R}_2$ ,

## 17

R<sub>5</sub>, R<sub>6</sub>=—H. Ten millimoles (1.80 grams) of caffeic acid was dissolved in 300 milliliters of 0.1 Normal aqueous sodium hydroxide (NaOH). The remaining procedure followed that of Example 1. The yield of synthetic soil extract was 0.65 gram.

## EXAMPLE 6

## Preparation of a Synthetic Humic Acid from Tetrahydroxybenzoquinone

The chemical structure of the starting organic compound is shown in Table 2. Ten millimoles (1.72 grams) of tetrahydroxybenzoquinone was dissolved in 300 milliliters of 0.1 Normal aqueous sodium hydroxide (NaOH). The remaining procedure followed that of Example 1. The yield of synthetic soil extract was 0.016 gram.

## EXAMPLE 7

## Preparation of a Synthetic Humic Acid from 1,4-dihydroxybenzene (Hydroquinone)

The starting organic compound is shown in Table 1, and represented by R<sub>1</sub>, R<sub>4</sub>=—OH; and R<sub>2</sub>, R<sub>3</sub>, R<sub>5</sub>, R<sub>6</sub>=—H. Ten millimoles (1.10 grams) of hydroquinone was dissolved in 300 milliliters of 0.1 Normal aqueous sodium hydroxide (NaOH). The remaining procedure followed that of Example 1. The yield of synthetic soil extract was 0.16 gram.

## EXAMPLE 8

## Preparation of a Synthetic Humic Acid from 3,4,5-trihydroxybenzoic Acid (Gallic Acid)

The starting organic compound is shown in Table 1, and represented by R<sub>1</sub>=—CH<sub>2</sub>CO<sub>2</sub>H; R<sub>3</sub>, R<sub>4</sub>, R<sub>5</sub>=—OH; and R<sub>2</sub>, R<sub>6</sub>=—H. Ten millimoles (1.70 grams) of gallic acid was dissolved in 300 milliliters of 0.1 Normal aqueous sodium hydroxide (NaOH). The remaining procedure followed that of Example 1. The yield of synthetic soil extract was 0.10 gram.

## EXAMPLE 9

## Preparation of a Synthetic Humic Acid from 2,5-dihydroxyphenylacetic Acid (Homogentisic Acid)

The starting organic compound is shown in Table 1, and consisted of R<sub>1</sub>=—CH<sub>2</sub>CO<sub>2</sub>H; R<sub>2</sub>, R<sub>5</sub>=—OH; and R<sub>3</sub>, R<sub>4</sub>, R<sub>6</sub>=—H. Ten millimoles (1.68 grams) of homogentisic acid was dissolved in 300 milliliters of 0.1 Normal aqueous sodium hydroxide (NaOH). The remaining procedure followed that of Example 1. The yield of synthetic soil extract was 0.20 gram.

The following Examples 10–13 are illustrative of the entire process of the preferred embodiments including step e), addition of salts. Addition of salts increases the activity of the synthetic humic acids produced by the method. Examples 10–13 illustrate that the synthetic humic acid materials produced according to the chemical processes and separation and isolation procedures of the preferred embodiments exhibit the physicochemical properties and attributes of typical naturally-occurring commercially-available humic acids and other soil extracts. Examples 10–13 also illustrate that the therapeutic indications of the synthetic humic acids produced according to the chemical processes and separation and isolation procedures of the preferred embodiments are those of soil extracts and humic acids in general, that is to say for viral-related disorders and diseases.

## 18

## EXAMPLE 10

## Preparation of Another Synthetic Humic Acid from 2,5-dihydroxyphenylacetic Acid (Homogentisic Acid)

The starting organic compound is shown in Table 1, and represented by R<sub>1</sub>=—CH<sub>2</sub>CO<sub>2</sub>H; R<sub>2</sub>, R<sub>5</sub>=—OH; and R<sub>3</sub>, R<sub>4</sub>, R<sub>6</sub>=—H. Six millimoles (1 gram) of homogentisic acid was dissolved in 300 milliliters of 0.1 Normal aqueous sodium hydroxide (NaOH). The solution pH was adjusted to 8.5 with 6 Normal HCl. One and one-half millimoles (0.32 gram) of sodium periodate (NaIO<sub>4</sub>) and 0.12 gram of sodium sulfide nonahydrate (Na<sub>2</sub>S·9H<sub>2</sub>O; 0.5 millimole) were added, and the solution was placed in a water-bath at 50° C. overnight. One milligram of boric acid (H<sub>3</sub>BO<sub>3</sub>; 0.016 millimole), 0.021 gram of ferrous sulfate heptahydrate (FeSO<sub>4</sub>·7H<sub>2</sub>O; 0.075 millimole), and 0.006 gram of calcium sulfate dihydrate (CaSO<sub>4</sub>·2H<sub>2</sub>O; 0.035 millimole) were added and the solution was stirred for 2 hours at room temperature. Any precipitate was removed by centrifugation. The solution was dialyzed with a 1,000-dalton cut-off flow-through open-channel or screen membrane system (Pall Filtron: Ultrasette® 7 Tangential Flow Device or Mini-Ultrasette® 7 Tangential Flow Device used with a Pall Filtron Ultralab® 7 Specialized Pump and Reservoir System) to a conductivity of 30 microsiemens or less against distilled water. The dialysis apparatus was then used to concentrate the solution to about 200 milliliters. The solution can be saved at this point for further use as an aqueous solution; or it can be freeze-dried to a powder. (Fifty to two hundred milligrams of mannose or other suitable carbohydrate can be added to the solution prior to freeze-drying to reduce static electricity effects associated with the freeze-dried powder.) The yield of synthetic soil extract was 0.23 gram.

The HPLC trace of the synthetic soil extract obtained in this Example is illustrated in FIG. 1. Peaks 1–6 were produced by this example. Peak 5 is under the shoulder of Peak 4 and is not overtly apparent. A mathematical first derivative of the detector signal versus time can more clearly show Peak 5. FIG. 2 shows the HPLC trace of a typical commercially-available natural humic acid. Peak 6 in FIGS. 1 and 2 was produced by a column wash with 90–100% v/v methanol and also contains synthetic humic acid. It can be seen that with the exception of the relative amounts of material in Peaks 2, 4, and 6, the remainder of the HPLC traces in FIGS. 1 and 2 are essentially equivalent. Thus, the synthetic procedure of the present invention produced a humic acid material with physicochemical characteristics that are essentially equivalent to those of a commercially-available soil extract.

## EXAMPLE 11

## Preparation of Still Another Synthetic Humic Acid from 2,5-dihydroxyphenylacetic Acid (Homogentisic Acid)

The starting organic compound is shown in Table 1, and represented by R<sub>1</sub>=—CH<sub>2</sub>CO<sub>2</sub>H; R<sub>2</sub>, R<sub>5</sub>=—OH; and R<sub>3</sub>, R<sub>4</sub>, R<sub>6</sub>=—H. Ten millimoles (1.68 grams) of homogentisic acid was dissolved in 300 milliliters of 0.1 Normal aqueous sodium hydroxide (NaOH). The remaining procedure followed that of the preceding Example. The yield of synthetic soil extract was 0.47 gram. The HPLC trace of the synthetic soil extract obtained in this Example was identical to that described in Example 10 and illustrated in FIG. 1.

EXAMPLE 12

Preparation of Still Another Synthetic Humic Acid from 2,5-Dihydroxyphenylacetic Acid (Homogentisic Acid)

The starting organic compound is shown in Table 1, and represented by R<sub>1</sub>=—CH<sub>2</sub>CO<sub>2</sub>H; R<sub>2</sub>, R<sub>5</sub>=—OH; and R<sub>3</sub>, R<sub>4</sub>, R<sub>6</sub>=—H. Ten millimoles (1.68 grams) of homogentisic acid was dissolved in 300 milliliters of 0.1 Normal aqueous sodium hydroxide (NaOH). The remaining procedure followed that of the preceding Example. The yield of synthetic soil extract was 0.4 gram. The HPLC trace of the synthetic soil extract obtained in this Example was identical to that described in Example 10 and illustrated in FIG. 1.

EXAMPLE 13

Preparation of Another Synthetic Humic Acid from 3,4-dihydroxycinnamic Acid (Cafeic Acid)

The starting organic compound is shown in Table 1, and represented by R<sub>1</sub>=CHCHCO<sub>2</sub>H; R<sub>3</sub>, R<sub>4</sub>=—OH; and R<sub>2</sub>, R<sub>5</sub>, R<sub>6</sub>=—H. Ten millimoles (1.8 grams) of homogentisic acid was dissolved in 300 milliliters of 0.1 Normal aqueous sodium hydroxide (NaOH). The remaining procedure followed that of the preceding Example. The yield of synthetic soil extract was 0.51 gram.

EXAMPLE 14

Preparation of a Synthetic Humic Acid from 1,3,4,5-tetrahydroxycyclohexane-carboxylic Acid 3-(3,4-dihydroxycinnamate) (Chlorogenic Acid)

The starting organic compound is shown in Table 2. Ten millimoles (3.54 grams) of chlorogenic acid was dissolved in 300 milliliters of 0.1 Normal aqueous sodium hydroxide (NaOH). The remaining procedure followed that of the preceding Example. The yield of synthetic soil extract was 0.23 gram.

EXAMPLE 15

In vitro Toxicity of Synthetic Humic Acid Prepared According to Examples 10, 11 and 12

Humic acid synthesized from homogentisic acid was prepared according to the procedure of Examples 10, 11 and 12. The in vitro toxicity of the materials was assessed as follows:

Five units of 450 milliliters each of whole human blood were collected into CP2D/AS-3 Leukotrap RC-PL systems. The blood was rested for 3 hours at room temperature. Each sample was weighed, and then centrifuged at 2820 revolutions per minute (2312 gravities) for 3 minutes, 44 seconds. The blood samples were then expressed through ATS-LPL filters into platelet storage bags. The filtration time was noted. The LR-PRP was centrifuged at 3600 revolutions per minute (3768 gravities) for 7 minutes. All but about 55 grams of platelet poor plasma was removed from each sample. The platelet concentrates were rested for 90 minutes at room temperature, and were then weighed and placed in a platelet incubator. RCM1 filters were primed with AS-3 solution. The primary bags were hung at a height of 60 inches above empty AS-3 bags, such that filtration occurred by gravity. The filtration time was noted, and the LRRCC systems were sealed off 3 inches below the RCM1 filters. Each RCM1 filter together with 6 inches of tubing and the LR-RCC, including the donor identification tube segment, were weighed. Samples were taken at this point for post-filtration testing (LR-RCC).

At Day 1 sufficient synthetic humic acid was added to each platelet concentrate so as to make its concentration 25 micrograms per milliliter. Treated platelet concentrates were then incubated in a platelet incubator for 1 hour, following which samples of each platelet concentrate were taken for testing. Subsequent samples were also taken on day 5 for further testing.

Table 3 shows the effect of the synthetic humic acid prepared as described in Example 10 on the viability of platelet concentrates as measured according to the procedures of this Example. The results were all nominal, that is, the synthetic humic acid had no effect on platelet viability (i.e., is nontoxic). The same results were obtained when the concentration of humic acid was made 100 micrograms per milliliter instead of 25 micrograms per milliliter. These results are particularly noteworthy, as blood platelets are known to be sensitive to a variety of chemical agents. It is for this reason that few safe anti-viral treatments are available for blood platelets.

TABLE 3

Unit No.	pH at 22° C.		pCO <sub>2</sub> , mm Hg		pO <sub>2</sub> , mm Hg		HCO <sub>3</sub> , mmol/L		MPV, fl	
	Day 1	Day 5	Day 1	Day 5	Day 1	Day 5	Day 1	Day 5	Day 1	Day 5
1	7.466	7.394	19.3	12.8	33.5	44.4	16.8	9.5	7.0	6.6
2	7.321	7.215	21.6	14.3	9.9	22.2	13.8	7.3	6.7	6.3
3	7.320	7.276	24.4	16.6	10.3	21.3	15.6	9.7	6.7	6.5
4	7.368	7.308	20.7	14.3	13.4	22.2	14.6	8.9	6.5	6.3
5	7.457	7.454	20.1	13.8	23.7	29.0	17.1	11.6	7.7	7.4
Mean	7.386	7.329	21.2	14.4	18.2	27.8	15.6	9.4	6.9	6.6
Std. Dev.	0.071	0.095	2.0	1.4	10.2	9.8	1.4	1.5	0.5	0.6

Unit No.	WBC Yield, ×10 <sup>5</sup>		Platelet Yield, ×10 <sup>10</sup>		Streaming		% ESC		% HSR		Lactate, mmol/L	
	Day 1	Day 5	Day 1	Day 5	Day 1	Day 5	Day 1	Day 5	Day 1	Day 5	Day 1	Day 5
1	0.1	8.3	9.0	3	3	24.2	16.9	78.0	64.0	5.1	12.1	
2	0.2	14.5	14.2	3	3	27.5	20.3	81.7	71.5	6.6	13.4	
3	0.4	13.3	13.4	3	3	28.7	26.3	81.7	79.4	6.3	12.4	



TABLE 3-continued

4	0.3	11.7	12.3	3	2	22.1	19.2	81.4	77.1	6.6	13.1
5	0.3	8.9	9.1	3	3	19.1	14.4	74.7	70.2	4.5	9.7
Mean	0.3	11.3	11.6	3.0	2.8	24.3	19.4	79.5	72.4	5.8	12.1
Std. Dev.	0.1	2.7	2.4	0.0	0.4	3.9	4.5	3.1	6.1	1.0	1.4

## EXAMPLE 16

In vitro Toxicity of Synthetic Humic Acid Prepared  
According to Examples 10–14

Humic acid synthesized from homogentisic acid was prepared according to the procedure of Examples 10–12. Humic acid synthesized from caffeic acid was prepared according to the procedure of Example 13. Humic acid synthesized from chlorogenic acid was prepared according to the procedure of Example 14. Natural-product humic acid was prepared by dialysis with subsequent freeze-drying as described in Examples 1–14. The in vitro toxicity of the materials was assessed as follows:

Cytotoxicity was examined with six concentrations of each humate material, and one “no-drug” concentration. All materials were tested in African green monkey kidney cells (CV-1; Diagnostic Hybrids, Inc., Athens, Ga.) in triplicate. The cells were provided in flat dishes containing multiple cell wells. The cells were cultured in the presence of different concentrations of humate materials for 24–36 hours at 35–37° C. in a CO<sub>2</sub>-humidified incubator. The morphology of the cultured cells was examined visually to determine any cytotoxic effects. No abnormal cell morphology was observed in cultures with “no drug” nor in any containing humate concentrations up to 500 micrograms per milliliter. Furthermore, no apparent CV-1 cell death (that is, cell detachment from the bottom of the wells) was observed at any concentration of any material tested. The results established that the materials were not cytotoxic at concentrations up to at least 500 micrograms per milliliter.

## EXAMPLE 17

In vivo Toxicity of Synthetic Humic Acid Prepared  
According to Examples 10–14

Humic acid synthesized from homogentisic acid was prepared according to the procedure of Examples 10–12. Humic acid synthesized from caffeic acid was prepared according to the procedure of Example 13. Humic acid synthesized from chlorogenic acid was prepared according to the procedure of Example 14. Natural-product humic acid was prepared by dialysis with subsequent freeze-drying as described in Examples 1–14.

The in vivo acute intravenous systemic toxicity of the humate materials was assessed as follows. Each humate material was dissolved separately in sterile, pyrogen-free 0.9% aqueous sodium chloride solvent to yield solutions of final concentrations of 1, 0.5 and 0.25 milligrams per milliliter. The test animals were viral antibody-free Swiss Webster mice, which weighed in the range of 17–23 grams at the time of testing. All test animals were quarantined and checked for signs of disease prior to testing. All test animals were group-housed five per cage in plastic cages with stainless steel suspended lids.

For each dose of each humate material, ten mice (five males and five females) were administered the sample humate material intravenously in the amount of 50 milliliters per kilogram body weight. Ten additional mice were simi-

larly administered 0.9% sodium chloride solution (the solvent vehicle) as a zero control. This procedure resulted in humate doses of 50 milligrams per kilogram body weight from the 1 milligram per milliliter solution, 25 milligrams per kilogram from the 0.5 milligram per milliliter solution, 12.5 milligrams per kilogram from the 0.25 milligram per milliliter solution, and 0 milligrams per kilogram from the 0.9% sodium chloride (blank) solution.

Following injection, the mice were offered a balanced Teklad diet and water ad libitum for the duration of the study. All mice were examined for viability for fourteen days. Zero time, day seven and day fourteen weights and toxic symptoms were recorded.

No mortalities were observed for any of the mice over the fourteen day observation period and, while some clinical findings were observed, they were not indicative of toxicity.

## EXAMPLE 18

Influenza A Cytoprotection Properties of Natural-  
Product and Synthetic Humic Acids Prepared  
According to Examples 10–14

Humic acid synthesized from homogentisic acid was prepared according to the procedure of Examples 10–12. Humic acid synthesized from caffeic acid was prepared according to the procedure of Example 13. Humic acid synthesized from chlorogenic acid was prepared according to the procedure of Example 14. Natural-product humic acid was prepared by dialysis with subsequent freeze-drying as described in Examples 1–14. The anti-viral properties of the humate materials were assessed according to the following methods:

Influenza virus type A (Beijing/262/95) (H1N1) was obtained from the Center for Disease Control (CDC). Influenza virus type A (Panama/2007/99) (H3N2) was also obtained from CDC. The test cell line consisted of Madin Darby canine kidney (MDCK) cells.

The first protocol used to determine the cytoprotection properties of the humate materials was as follows: each test was run in 96-well flat-bottomed microplates. Four log<sub>10</sub> dilutions of each test compound were added to 3 cups containing the cell monolayer. Within 5 minutes the virus was added and the plate sealed, incubated at 37° C., and the drug inhibition of the viral cytopathic effect (CPE) assessed microscopically when untreated infected controls had developed a 3 to 4+ CPE (approximately 72 to 120 hours). The data were expressed as 50% effective inhibitory concentration (CPE EC<sub>50</sub>).

The second cytoprotection assay protocol, employed to validate the CPE results, was as follows: in the neutral red dye uptake version of the antiviral test protocol described immediately above, the same 96-well microplates were utilized after the CPE had been read. Neutral red dye was then added to the medium: cells not damaged by virus take up a greater amount of dye, which can be read on a computerized micro plate autoreader (the method is described in detail by McManus: *Appl. Environ. Microbiol.* 1976, 31, 35–38). The data were expressed as 50% effective inhibitory concentration (NR EC<sub>50</sub>).

The results of drug efficacy testing against influenza A viruses with the above-described protocols are summarized in Table 4.

This Example demonstrates that the humate materials can protect cells against infection by influenza A viruses.

TABLE 4

Compound <sup>1</sup>	Influenza A (H1N1)		Influenza A (H3N2)	
	CPE EC <sub>50</sub>	NR EC <sub>99</sub>	CPE EC <sub>50</sub>	NR EC <sub>99</sub>
Caffeic Acid Humate	1	0.6	<1	<1
Chlorogenic Acid Humate	45	40	6	6.5
Homogentisic Acid Humate	3.7	3.2	4.5	3.2
Natural-Product Humate	2.5	2.5	<1	<1

<sup>1</sup>All concentrations in micrograms per milliliter.

## EXAMPLE 19

Influenza B Cytoprotection Properties of Natural-Product and Synthetic Humic Acids Prepared According to Examples 10–14

Humic acid synthesized from homogentisic acid was prepared according to the procedure of Examples 10–12. Humic acid synthesized from caffeic acid was prepared according to the procedure of Example 13. Humic acid synthesized from chlorogenic acid was prepared according to the procedure of Example 14. Natural-product humic acid was prepared by dialysis with subsequent freeze-drying as described in Examples 1–14. The anti-viral properties of the humate materials were assessed according to the following methods:

Influenza virus type B (Beijing/184/93) (H1N1) was obtained from the Center for Disease Control (CDC). The test cell line consisted of Madin Darby canine kidney (MDCK) cells.

The two protocols used to determine the cytoprotection properties of the humate materials were as described in Example 18.

The results of drug efficacy testing against influenza B viruses with the above-described protocols are summarized in Table 5.

This Example demonstrates that the humate materials can protect cells against infection by influenza B viruses.

TABLE 5

Compound <sup>1</sup>	Influenza B	
	CPE EC <sub>50</sub>	NR EC <sub>99</sub>
Caffeic Acid Humate	<1	<1
Chlorogenic Acid Humate	5.5	4.7
Homogentisic Acid Humate	3.2	3.2
Natural-Product Humate	<1	<1

<sup>1</sup>All concentrations in micrograms per milliliter.

The Examples described above establish relevant influenza antiviral data and efficacy of the synthetic humate compounds. The scope of the studies conforms with current requirements put forth by the FDA for pre-clinical analysis of new anti-virals.

Various modifications and alterations of this invention will be apparent to those skilled in the art without departing from the scope and spirit of this invention. It should be

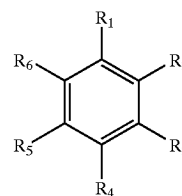
understood that the invention is not limited to the embodiments disclosed therein, and that the claims should be interpreted as broadly as the prior art allows.

What is claimed is:

1. A method for inhibiting and/or treating influenza virus infections in a mammal comprising administering to the mammal an effective amount of a synthetic phenolic polymeric material, said synthetic phenolic material having been prepared by the following steps:

- dissolving in an aqueous solution at least one starting organic compound comprising at least one hydroxyl group and at least one carbonyl group or at least two hydroxyl groups on an aromatic structure;
- adjusting the pH of the aqueous solution resulting from step a) to between about 8 and 11;
- oxidizing the at least one starting organic compound solution resulting from step b);
- polymerizing the oxidized compound from step c);
- adding at least one water soluble compound or salt selected from the group consisting of boric acid, borate salts, alkaline earth salts, transition metal salts, alkaline sulfides, alkaline earth sulfides, or transition metal sulfides to the aqueous solution resulting from step d); and
- removing molecules from the solution resulting from step e) below about 500 to about 10,000 daltons.

2. The method according to claim 1, wherein the starting organic compound is selected from the group consisting of a compound represented by the Formula I:

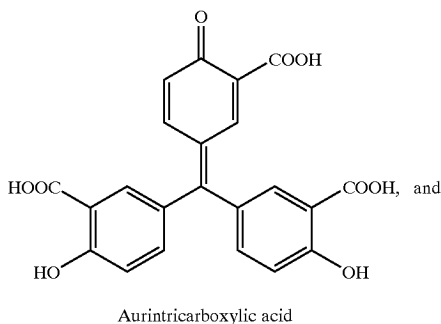
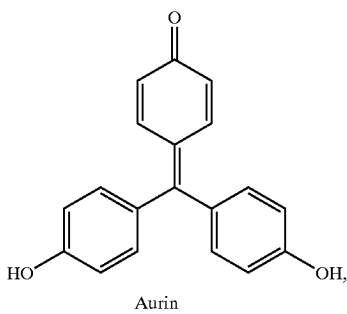
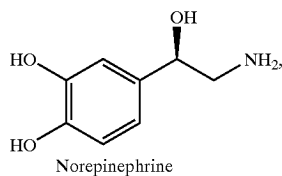
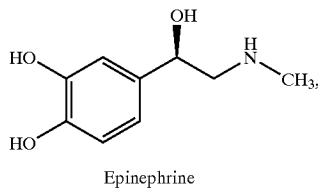
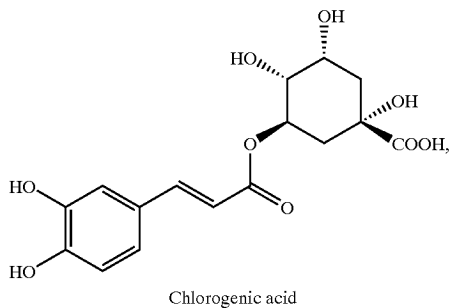
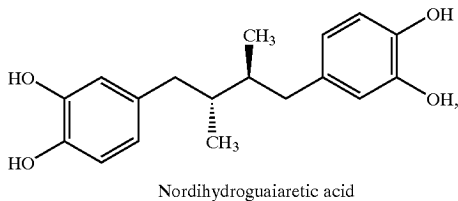


wherein R<sub>1</sub>, R<sub>2</sub>, R<sub>3</sub>, R<sub>4</sub>, R<sub>5</sub>, and R<sub>6</sub> is a substituent selected from the group consisting of H, CH<sub>3</sub>, CH<sub>2</sub>CH<sub>3</sub>, (CH<sub>2</sub>)<sub>2</sub>CH<sub>3</sub>, CH(CH<sub>3</sub>)<sub>2</sub>, OH, OCH<sub>3</sub>, CHO, CO<sub>2</sub>H, CO<sub>2</sub>CH<sub>3</sub>, CH<sub>2</sub>OH, CH<sub>2</sub>OCH<sub>3</sub>, CH<sub>2</sub>CHO, CH<sub>2</sub>CO<sub>2</sub>H, CH<sub>2</sub>CO<sub>2</sub>CH<sub>3</sub>, (CH<sub>2</sub>)<sub>2</sub>OH, (CH<sub>2</sub>)<sub>2</sub>OCH<sub>3</sub>, (CH<sub>2</sub>)<sub>2</sub>CHO, (CH<sub>2</sub>)<sub>2</sub>CO<sub>2</sub>H, (CH<sub>2</sub>)<sub>2</sub>CO<sub>2</sub>CH<sub>3</sub>, CH(CH<sub>3</sub>)OH, CH(CH<sub>3</sub>)OCH<sub>3</sub>, CH(CH<sub>3</sub>)CHO, CH(CH<sub>3</sub>)CO<sub>2</sub>H, CH(CH<sub>3</sub>)CO<sub>2</sub>CH<sub>3</sub>, CH(CH<sub>3</sub>)CH<sub>2</sub>OH, CH(CH<sub>3</sub>)CH<sub>2</sub>OCH<sub>3</sub>, CH(CH<sub>3</sub>)CH<sub>2</sub>CHO, CH(CH<sub>3</sub>)CH<sub>2</sub>CO<sub>2</sub>H, CH(CH<sub>3</sub>)CH<sub>2</sub>CO<sub>2</sub>CH<sub>3</sub>, CH(OH)<sub>2</sub>, CH(OH)OCH<sub>3</sub>, CH(OH)CHO, CH(OH)CO<sub>2</sub>H, CH(OH)CO<sub>2</sub>CH<sub>3</sub>, CH(OCH<sub>3</sub>)OH, CH(OCH<sub>3</sub>)<sub>2</sub>, CH(OCH<sub>3</sub>)CHO, CH(OCH<sub>3</sub>)CO<sub>2</sub>H, CH(OCH<sub>3</sub>)CO<sub>2</sub>CH<sub>3</sub>, CH(OH)CH<sub>2</sub>OH, CH(OH)CH<sub>2</sub>OCH<sub>3</sub>, CH(OH)CH<sub>2</sub>CHO, CH(OH)CH<sub>2</sub>CO<sub>2</sub>H, CH(OH)CH<sub>2</sub>CO<sub>2</sub>CH<sub>3</sub>, CH(OCH<sub>3</sub>)CH<sub>2</sub>OH, CH(OCH<sub>3</sub>)CH<sub>2</sub>OCH<sub>3</sub>, CH(OCH<sub>3</sub>)CH<sub>2</sub>CHO, CH(OCH<sub>3</sub>)CH<sub>2</sub>CO<sub>2</sub>H, CH(OCH<sub>3</sub>)CH<sub>2</sub>CO<sub>2</sub>CH<sub>3</sub>, (CH<sub>2</sub>)<sub>3</sub>OH, —(CH<sub>2</sub>)<sub>3</sub>OCH<sub>3</sub>, (CH<sub>2</sub>)<sub>3</sub>CHO, (CH<sub>2</sub>)<sub>3</sub>CO<sub>2</sub>H, (CH<sub>2</sub>)<sub>3</sub>CO<sub>2</sub>CH<sub>3</sub>, CHCHOH (cis or trans), CHCHOCH<sub>3</sub> (cis or trans), CHCHCHO (cis or trans), CHCHCO<sub>2</sub>H (cis or trans), CHCHCO<sub>2</sub>CH<sub>3</sub> (cis or trans), CH<sub>2</sub>CHCHOH (cis or trans), CH<sub>2</sub>CHCHOCH<sub>3</sub> (cis or trans), CH<sub>2</sub>CHCHCHO (cis or trans), CH<sub>2</sub>CHCHCO<sub>2</sub>H (cis or trans), and CH<sub>2</sub>CHCHCO<sub>2</sub>CH<sub>3</sub> (cis or trans).

3. The method according to claim 2, wherein the compound comprises at least one hydroxyl group and at least one carboxylic acid group.

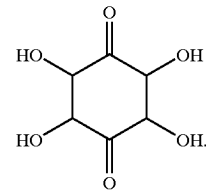
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4. The method according to claim 1, wherein the starting organic compound is selected from the group consisting of



26

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Tetrahydroxybenzoquinone

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5. The method according to claim 1, wherein the aqueous solution in step a) comprises sodium hydroxide.

6. The method according to claim 1, wherein the temperature in step d) is between about 35° C. and 80° C.

7. The method according to claim 1, wherein the temperature in step f) is between about 20° C. and 80° C.

8. The method according to claim 1, wherein the method of preparation of the synthetic phenolic polymeric material further comprises a step, following step g), of:

h) concentrating the solution resulting from step g).

9. The method according to claim 8, wherein the method of preparation of the synthetic phenolic polymeric material further comprises a step, following step h), of:

i) removing the water from the solution resulting from the step h).

10. The method according to claim 1, wherein the influenza virus infection is effected by a virus selected from the group consisting of influenza A and influenza B.

11. The method according to claim 1, wherein the mammal is a human.

12. The method according to claim 1, wherein administering the effective amount of synthetic phenolic polymeric material is performed systemically.

13. The method according to claim 1, wherein administering the effective amount of synthetic phenolic polymeric material is performed topically.

14. The method according to claim 1, further comprising administering an additional antiviral composition in combination with the effective amount of a synthetic phenolic polymeric material.

15. The method according to claim 14, wherein administering the effective amount of synthetic phenolic polymeric material and the antiviral composition is performed systemically.

16. The method according to claim 14, wherein administering the effective amount of synthetic phenolic polymeric material and the antiviral composition is performed topically.

17. A method of inhibiting influenza viral attachment to host cells in a mammal comprising administering to the mammal an effective amount of a synthetic phenolic polymeric material, said synthetic phenolic material having been prepared by the following steps:

a) dissolving in an aqueous solution at least one starting organic compound comprising at least one hydroxyl group and at least one carbonyl group or at least two hydroxyl groups on an aromatic structure;

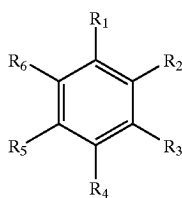
b) adjusting the pH of the aqueous solution resulting from step a) to between about 8 and 11;

c) oxidizing the at least one starting organic compound solution resulting from step b);

27

- d) polymerizing the oxidized compound from step c);  
 e) adding at least one water soluble compound or salt selected from the group consisting of boric acid, borate salts, alkaline earth salts, transition metal salts, alkaline sulfides, alkaline earth sulfides, or transition metal sulfides to the aqueous solution resulting from step d); and  
 f) removing molecules from the solution resulting from step e below about 500 to about 10,000 daltons.

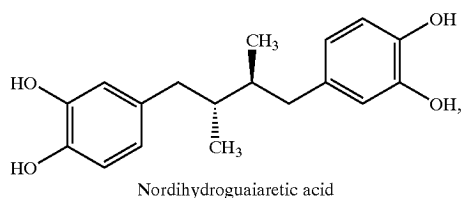
18. The method according to claim 17, wherein the starting organic compound is selected from the group consisting of a compound represented by the formula I:



wherein  $R_1$ ,  $R_2$ ,  $R_3$ ,  $R_4$ ,  $R_5$ , and  $R_6$  is a substituent selected from the group consisting of H,  $CH_3$ ,  $CH_2CH_3$ ,  $(CH_2)_2CH_3$ ,  $CH(CH_3)_2$ , OH,  $OCH_3$ , CHO,  $CO_2H$ ,  $CO_2CH_3$ ,  $CH_2OH$ ,  $CH_2OCH_3$ ,  $CH_2CHO$ ,  $CH_2CO_2H$ ,  $CH_2CO_2CH_3$ ,  $(CH_2)_2OH$ ,  $(CH_2)_2OCH_3$ ,  $(CH_2)_2CHO$ ,  $(CH_2)_2CO_2H$ ,  $(CH_2)_2CO_2CH_3$ ,  $CH(CH_3)OH$ ,  $CH(CH_3)OCH_3$ ,  $CH(CH_3)CHO$ ,  $CH(CH_3)CO_2H$ ,  $CH(CH_3)CO_2CH_3$ ,  $CH(CH_3)CH_2OH$ ,  $CH(CH_3)CH_2CO_2H$ ,  $CH(CH_3)CH_2CO_2CH_3$ ,  $CH(CH_3)CH_2CHO$ ,  $CH(CH_3)CH_2CO_2H$ ,  $CH(CH_3)CH_2CO_2CH_3$ ,  $CH(OH)_2$ ,  $CH(OH)OCH_3$ ,  $CH(OH)CHO$ ,  $CH(OH)CO_2H$ ,  $CH(OH)CO_2CH_3$ ,  $CH(OCH_3)OH$ ,  $CH(OCH_3)_2$ ,  $CH(OCH_3)CHO$ ,  $CH(OCH_3)CO_2H$ ,  $CH(OCH_3)CO_2CH_3$ ,  $CH(OH)CH_2OH$ ,  $CH(OH)CH_2OCH_3$ ,  $CH(OH)CH_2CHO$ ,  $CH(OH)CH_2CO_2H$ ,  $CH(OH)CH_2CO_2CH_3$ ,  $CH(OCH_3)CH_2OH$ ,  $CH(OCH_3)CH_2OCH_3$ ,  $CH(OCH_3)CH_2CHO$ ,  $CH(OCH_3)CH_2CO_2H$ ,  $CH(OCH_3)CH_2CO_2CH_3$ ,  $(CH_2)_3OH$ ,  $-(CH_2)_3OCH_3$ ,  $(CH_2)_3CHO$ ,  $(CH_2)_3CO_2H$ ,  $(CH_2)_3CO_2CH_3$ ,  $CHCHOH$  (cis or trans),  $CHCHOCH_3$  (cis or trans),  $CHCHCHO$  (cis or trans),  $CHCHCO_2H$  (cis or trans),  $CHCHCO_2CH_3$  (cis or trans),  $CH_2CHCHOH$  (cis or trans),  $CH_2CHCHOCH_3$  (cis or trans),  $CH_2CHCHCHO$  (cis or trans),  $CH_2CHCHCO_2H$  (cis or trans), and  $CH_2CHCHCO_2CH_3$  (cis or trans).

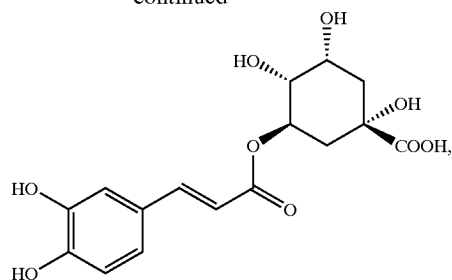
19. The method according to claim 18, wherein the compound comprises at least one hydroxyl group and at least one carboxylic acid group.

20. The method according to claim 17, wherein the starting organic compound is selected from the group consisting of

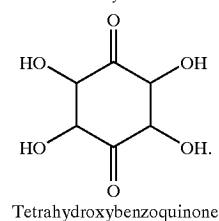
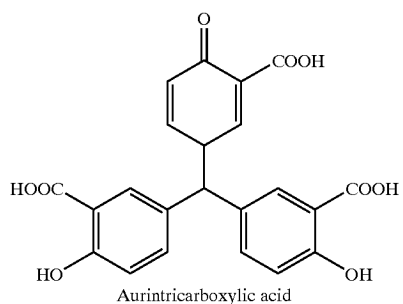
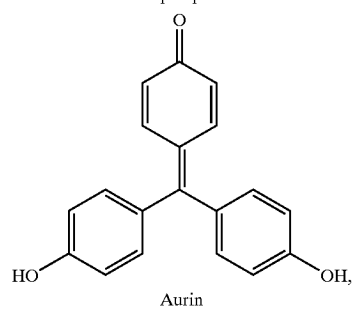
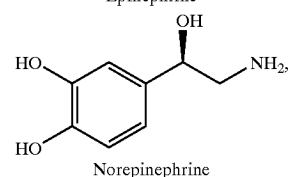
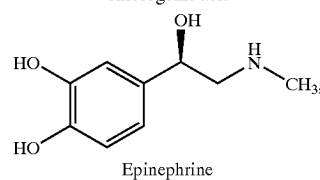


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(I)



21. The method according to claim 17, wherein the aqueous solution in step a) comprises sodium hydroxide.

22. The method according to claim 17, wherein the temperature in step d) is between about 35° C. and 80° C.

23. The method according to claim 17, wherein the temperature in step f) is between about 20° C. and 80° C.

24. The method according to claim 17, wherein the method of preparation of the synthetic phenolic polymeric material further comprises a step, following step g), of:

h) concentrating the solution resulting from step g).

25. The method according to claim 24, wherein the method of preparation of the synthetic phenolic polymeric material further comprises a step, following step h), of:

i) removing the water from the solution resulting from the step h).

26. The method according to claim 17, wherein the influenza virus infection is effected by a virus selected from the group consisting of influenza A and influenza B.

27. The method according to claim 17, wherein the mammal is a human.

28. The method according to claim 17, wherein administering the effective amount of synthetic phenolic polymeric material is performed systemically.

29. The method according to claim 17, wherein administering the effective amount of synthetic phenolic polymeric material is performed topically.

30. The method according to claim 17, further comprising administering an additional antiviral composition in combination with the effective amount of a synthetic phenolic polymeric material.

31. The method according to claim 30, wherein administering the effective amount of synthetic phenolic polymeric material and the antiviral composition is performed systemically.

32. The method according to claim 30, wherein administering the effective amount of synthetic phenolic polymeric material and the antiviral composition is performed topically.

33. The method according to claim 1, wherein the synthetic phenolic material has been prepared by a method wherein step c) comprises adding an alkaline periodate salt

or alkaline-earth periodate salt to the aqueous solution resulting from step b).

34. The method according to claim 1, wherein the synthetic phenolic material has been prepared by a method wherein step d) comprises maintaining the temperature of the solution from step c) between about 20° C. and 100° C. for a period of at least about 30 minutes.

35. The method according to claim 1, wherein the synthetic phenolic material has been prepared by a method wherein by a method further comprising allowing the aqueous solution from step e) to stand with or without stirring at about 20° C. to 100° C. for at least about 2 hours after step e).

36. The method according to claim 17, wherein the synthetic phenolic material has been prepared by a method wherein step c) comprises adding an alkaline periodate salt or alkaline-earth periodate salt to the aqueous solution resulting from step b).

37. The method according to claim 17, wherein the synthetic phenolic material has been prepared by a method wherein step d) comprises maintaining the temperature of the solution from step c) between about 20° C. and 100° C. for a period of at least about 30 minutes.

38. The method according to claim 17, wherein the synthetic phenolic material has been prepared by a method wherein by a method further comprising allowing the aqueous solution from step e) to stand with or without stirring at about 20° C. to 100° C. for at least about 2 hours after step e).

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